

5 METHOD OF MANUFACTURING CITRIC ACID CHELATES

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Field of the Invention

10 This invention relates to chelated metals having particular utility as plant micronutrients. More particularly, this invention relates to a method of manufacturing chelated metals for subsequent pre- or postemergent application to agricultural crops, either alone or in combination with other fertilizing material.

Background of the Invention

15 Micronutrients are those nutrients which are essential for plant growth but which are required in amounts which are typically less than that required for macronutrients such nitrogen, phosphorus and potassium. The seven micronutrients commonly understood to be necessary for plant growth are boron, chlorine, copper, iron, manganese, molybdenum and zinc. Cobalt, calcium, magnesium and sulfur may also be required in smaller amounts, and
20 are variously classified as micronutrients or secondary nutrients. As used herein, the term micronutrients includes secondary nutrients.

The constituent elements of micronutrients must be presented to plants in a form suitable for plant uptake, translocation and assimilation. One way in which certain micronutrient metals have been successfully applied to and
25 utilized by plants is by application of metal chelate solutions. The term chelate is from the Greek for claw, and describes a "complex" in which organic molecules form ring-like structures through coordinate bonds with metal ions. The metal chelate structure loosely ties up the constituent metal ion(s) in a form in which the metal component is less likely to dissociate or
30 participate in chemical reactions when the chelate is in the soil or in solution with other fertilizers. This allows the chelated metal to remain available until after the product penetrates and translocates within plant.

Chelated metals are currently available in the fertilizer industry in powder and liquid forms. Each form has certain disadvantages, however, and

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both forms tend to be more expensive than inorganic metal compounds. For example, aqueous metal chelate formulations typically are supplied in 2.5 or five gallon plastic jugs containing from 5% to 9% by weight of the chelated metal. Cost of shipping these products is high, primarily because of the relatively low weight percent of the chelated metal in the product as compared to the volume of water in which the metal chelate is dissolved. In addition, if temperatures drop below 32°F some metal chelates crystallize out of liquid formulations and concentrate at the bottom of the container. Naturally this creates problems when subsequently applying the metal chelate solution to crops.

Not only does the bulkiness of the containers pose a storage space problem for some growers or fertilizer dealers, disposal of the plastic jugs can be problematical because many landfills no longer accept these containers because of container bulkiness and contamination issues. To address problems associated with use of plastic jugs, growers crops for which metal chelate requirements are substantial may choose to buy aqueous metal chelates solutions in bulk and place the liquids in large tanks or other storage vessels. However, to avoid accidental spillage when using such storage vessels, individual storage vessels are preferably placed in a lined dike having a capacity 110 % of the volume of the largest storage vessel. This technique involves substantial cost to the grower or fertilizer dealer, and does not obviate crystallization problems of metal chelate solutions at lower temperatures.

While powder metal chelates are available and solve some of the stability, storage and disposal problems associated with metal chelate solutions, other problems remain unaddressed. For example, to produce certain powdered metal chelates, constituents are reacted in an aqueous solution and then dried, a process requiring substantial time, energy and expense. Powdered metal chelates produced in this manner include ethylene diamine tetraacetic acid (EDTA) chelates and lignin sulfonate- based metal chelates. While these powdered metal chelates can provide the constituent metal in greater concentration than metal chelate solutions, (e.g., 6% to 15% weight percent), powdered metal chelates are substantially more expensive,

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e.g., \$1.50 to \$8.00 per pound. The high cost of these products tends to limit their use to the horticulture industry.

Furthermore, some metal chelates solutions are decomposed by acid. The constituent metals may then exhibit reduced agronomic effectiveness.

5 Thus, mixing some metal chelate solutions, including reconstituted powdered metal chelates, with certain acidic fertilizers must be avoided prior to application of the metal chelate solution to plants. Yet application of micronutrients with macronutrient fertilizers is typically preferred, because it provides a way to more evenly distribute a small micronutrient volume over a
10 large target acreage.

It is against this background that the significant improvements and advancements of the present invention have taken place.

Objects of the Invention

15 It is the principal object of the present invention to provide agricultural micronutrients in a form which is economical to transport and in containers with minimal disposal problems.

It is another object of the present invention to provide agricultural micronutrients in a form which may be combined with other fertilizer materials to create a stable composition in which the agronomic effectiveness of the
20 micronutrients and the fertilizer material is maintained.

It is a further object of the present invention to manufacture agricultural micronutrients in a manner which does not require the removal of substantial amounts of water during processing and which has a shortened processing time over the processing time required to produce powdered EDTA metal
25 chelates.

Summary of the Invention

In accordance with its major aspects, a method of manufacturing a composition specially adapted for subsequent solubilizing and application to agricultural crops is disclosed. The method of the present invention includes
30 the steps of mixing a divalent metal salt, citric acid and sodium citrate to obtain a mixture having a moisture content most preferably of from approximately 10% to 1.25% by weight. The mixture is then processed in a drying environment to obtain a product which is partially chelated and which

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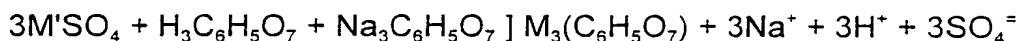
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(2) citric acid, and (3) sodium citrate, and the placement of this mixture in a drying environment. Prior to placement in the drying environment, the mixture preferably has a moisture content, by weight, of from approximately 10% to approximately 1.25%, with the most preferred moisture content, by weight of the total mixture, of from approximately 1.5% to approximately 1.25%. The primary reactants are preferably supplied to the mixture in molar ratios of 3:1:1 of divalent metal salt, citric acid, and sodium citrate, respectively. The preferred divalent metals are iron, copper, manganese and zinc. Suitable metal salts include sulfates, nitrates and phosphates.

The mixture is then heated in a controlled drying environment, producing a partially chelated product which has a total moisture content less than that of the mixture. The preferred moisture content of the partially chelated product, by weight, is from approximately 5% to 0.25%, with the most preferred moisture content of the final product, by weight, from approximately 1.25% to approximately 0.5%.

When a divalent metal sulfate is employed in the present invention, the equilibrium reaction established in the drying environment is:



Essentially, controlled dehydration and partial chelation of the mixture occurs in the drying environment, and a partially chelated citric acid product is produced thereby.

More particularly, Examples I-IV below describe the production of copper, zinc, iron and manganese citric acid products, respectively, which are partially chelated in accordance with the methods of the present invention. It should be noted that in addition to including the three primary reactants identified above, two secondary reactants are also added to the mixtures described in each example. Sodium tripolyphosphate is added as a wetting agent to the mixture in an amount equal to approximately 0.1% by weight of the final mixture. Monopotassium phosphate is added to condition the mixture, in an amount equal to approximately 2.5% by weight of the final mixture.

EXAMPLE I

1,600 pounds of copper sulfate salt (approximately 25% by weight

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copper), 180 pounds of citric acid dry soluble 80% active ingredient (AI), 180 pounds of sodium citrate dry soluble 80% AI, 4 pounds of dry sodium tripolyphosphate, and 36 pounds of dry monopotassium phosphate (0-52-34) were mixed, with the total moisture content of the mixture preferably from approximately 1.5% to approximately 1.25% by weight of the mixture. The copper sulfate mixture was then placed on a continuous flow, fluidized bed drying operated in a temperature range of from approximately 120°F to approximately 150°F, which constituted a drying environment for the mixture. The mixture was maintained in the drying environment for from approximately 50 seconds to approximately 70 seconds, during which time the moisture content of the resulting product was reduced to approximately 0.5% by weight of the product, and the product was partially chelated. The resulting product, which included copper citric acid chelate, was dry and friable. The product was allowed to cool and stabilize, and thereafter was introduced into a high speed mill to reduce particle size of the product and to reduce the tendency of the product to agglomerate. The ground product was then placed in containers, for example, plastic packages of 8 mil or greater thickness, and hermetically sealed. The ambient environment in which the product was maintained after treatment in the fluidized bed and before hermetic sealing had a relative humidity of less than 50%.

EXAMPLE II

1,475 pounds of zinc sulfate salt (approximately 35% by weight zinc), 243 pounds of citric acid dry soluble 80% AI, 242 pounds of sodium citrate dry soluble 80% AI, 6 pounds of dry sodium tripolyphosphate, and 34 pounds of dry monopotassium phosphate were mixed, with the total moisture content of the mixture preferably from approximately 1.5% to approximately 1.25% by weight of the mixture. The zinc sulfate mixture was then placed on a continuous flow, fluidized bed drying operated in a temperature range of from approximately 120°F to approximately 150°F. The mixture was maintained in this environment for approximately 50 seconds to 70 seconds, during which time the moisture content was reduced to approximately 0.5% by weight, and part of the zinc sulfate salt was reacted to form zinc citric acid chelate. The product was allowed to cool, and stabilize, and thereafter milled. The ground

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product was then placed in containers and hermetically sealed, as described above in Example I.

EXAMPLE III

1,350 pounds of iron sulfate salt (31% iron by weight), 200 pounds of
5 citric acid dry soluble 80% AI, 200 pounds of sodium citrate dry soluble 80%
AI, 6 pounds of dry sodium tripolyphosphate, and 244 pounds of dry
monopotassium phosphate were mixed, with the total moisture content of the
mixture preferably from approximately 1.5% to approximately 1.25% by
weight of the mixture. The iron sulfate mixture was then placed on a
10 continuous flow, fluidized bed drying operated in a temperature range of from
approximately 120°F to approximately 150°F. The mixture was maintained in
this environment for approximately 50 seconds to 70 seconds, during which
time the moisture content was reduced to approximately 0.5% by weight, and
part of the iron sulfate salt was reacted to form iron citric acid chelate. The
15 product was allowed to cool, and stabilize, and thereafter milled. The ground
product was then placed in containers and hermetically sealed, as described
above in Example I.

EXAMPLE IV

1,350 pounds of manganese sulfate salt (approximately 31% by weight
20 manganese), 200 pounds of citric acid dry soluble 80% AI, 200 pounds of
sodium citrate dry soluble 80% AI, 6 pounds of dry sodium tripolyphosphate,
and 244 pounds of dry monopotassium phosphate were mixed, with the total
moisture content of the mixture preferably from approximately 1.5% to
approximately 1.25% by weight of the mixture. The manganese sulfate
25 mixture was then placed on a continuous flow, fluidized bed drying operated
in a temperature range of from approximately 120°F to approximately 150°F.
The mixture was maintained in this environment for approximately 50
seconds to 70 seconds, during which time the moisture content was reduced
to approximately 0.5% by weight, and part of the manganese sulfate salt was
30 reacted to form manganese citric acid chelate. The product was allowed to
cool, and stabilize, and thereafter milled. The ground product was then
placed in containers and hermetically sealed, as described above in Example
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It has been discovered that the preferred products of the present invention are those manufactured in accordance with the procedures of Examples I-IV above, in which the mixture to be placed in the drying environment has a moisture content of 4% or less by weight. In contrast, when the finished product has a moisture content of 5% or more, the product tends to agglomerate after production and harden during storage. Product shelf life is reduced, and, in extreme cases, the hardened product cannot be solubilized and further chelation induced without regrinding the product. Also, the hardened product typically cannot be mixed with other fertilizer material without regrinding. Thus, it is preferred that the mixtures of the present invention have a total moisture content by weight not exceeding 4% and the final partially chelated product have a preferred moisture content by weight of from 0.25% to 3%, most preferably approximately 0.5%.

It has also been discovered that the preferred products of the present invention are those manufactured in accordance with the procedures of Examples I-IV above, in which the mixture to be placed in the drying environment having an ambient temperature of at least 101°F. Mixtures placed in a drying environment of 90°F to 100°F or less result in a partially chelated product which exhibits reduced solubility and reduced compatibility with some fertilizer materials. In addition, the product is susceptible to uncontrolled chemical reaction when the packaged product is stored in ambient temperature above 100°F, such as may be experienced in warehouses, truck trailers or other storage environments. Such uncontrolled chemical reactivity can result in destruction of product packaging and substantially reduced product shelf life. Thus, it is preferred that the mixtures of the present invention be treated in a drying environment of at least 101°F, preferably of from 101°F to 190°F, and most preferably from approximately 120°F to approximately 150°F.

It has also been determined that the manufacturing methods of the present invention have a limited degree of success when the mixtures described in Examples I-IV are maintained in a drying environment for at least 10 seconds and at most 30 minutes. It is preferred, however, that the mixtures be maintained in a drying environment for from 30 seconds to 5

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minutes, with maintenance of the mixtures in a drying environment for approximately 50 seconds to approximately 70 seconds most preferred. Mixtures maintained in a drying environment for an excessive or insufficient period of time may have inconsistent moisture content and texture, and have a resultant reduction in shelf life, solubility and fertilizer compatibility.

Examples V-VIII below describe the solubilization and further chelation of the concentrated, partially chelated ground products manufactured above in accordance with Examples I-IV, respectively, to produce metal chelate solutions suitable for pre- or postemergent application to agricultural crops. In each of Examples V-VIII, the addition of water to the ground product results in the solubilization of the ground product and further chelation of the divalent metal of the particular divalent metal salt present in the ground product.

EXAMPLE V

Two and one half pounds of the partially chelated copper/citric acid product produced according to Example I were added to one gallon of water. The product was mixed until the resulting solution was clear.

EXAMPLE VI

Sixteen pounds of the partially chelated zinc/citric acid product produced according to Example II were added to 3 gallons of water. The product was mixed until the resulting solution was clear.

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EXAMPLE VII

Two and one half pounds of the partially chelated iron/citric acid product produced according to Example III were added to 1 gallon of water. The product was mixed until the resulting solution was clear.

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EXAMPLE VIII

Two and one half pounds of the partially chelated manganese/citric acid product produced according to Example IV were added to 1 gallon of water. The product was mixed until the resulting solution was clear.

To ascertain the solubility and stability in fertilizer solutions of the metal citric acid chelate solutions produced above in accordance with the preferred methods of the present, 25 grams of a partially chelated copper/citric acid mixture prepared in accordance with Example I was further chelated and dissolved in 83 grams of water to produce a copper/citric acid chelate solution (referred to in Table I as the "Chelate Solution"). Thereafter, 18 grams of each of 13 different fertilizer solutions was separately mixed with 1.7 grams of the Chelate Solution, and the resulting solution visually evaluated upon mixing and after two hours if any undesirable precipitation or crystallization occurred or other reaction products were produced. The evaluation results are summarized in Table I.

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TABLE I

FERTILIZER SOLUTION	SOLUBILITY IN CHELATE SOLUTION INITIALLY UPON MIXING	SOLUBILITY IN CHELATE SOLUTION AFTER 2 HOURS
1. 25% urea solution (46-0-0)	Clear	Clear
2. 32% urea-ammonium nitrate solution	Clear	Clear
3. Potassium nitrate solution (13.8-0-46)	Clear	Clear
4. Ammonium polyphosphate (10-34-0)	Clear	Clear

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5. Potassium chloride solution (0-0-62)	Clear	Clear
6. Potassium hydroxide (0-0-75)	Dark flocculant	Top 95% of solution clear, with flocculant settled at bottom of container
7. Calcium chloride solution (10% Ca)	Fine white precipitate	Top 95% of solution clear, with precipitate settled at bottom of container
8. Phosphoric acid solution (0-0-80)	Clear	Clear
9. Aqua ammonia solution (24-0-0)	Clear blue solution	Clear blue solution
10. 2.5% P ₂ O ₅ -1.5% K ₂ O solution (0-15-0)	Insoluble at concentration	Insoluble at concentration
11. Ammonium thiosulfate (12-0-0 and 26 S)	Clear	Clear
12. 10% sodium lignosulfonate	Dark brown solution without precipitate or flocculant	Dark brown solution without precipitate or flocculant
13. Calcium nitrate (17-0-0 and 24 Ca)	Fine white precipitate	Top 95% of solution clear, with precipitate settled at bottom of container

By employing the methods of the present invention to manufacture a partially-chelated, citric acid product, complete solubilization and subsequent dehydration of the product during manufacture is not required. Substantial energy costs and the processing time associated with dehydration is successfully avoided. Indeed, by not first solubilizing the primary reactants to

achieve chelation and then dehydrating the fully chelated product, large volumes of water are not required at the manufacturing site. In addition, because the partially chelated, citric acid products of the present invention do not contain substantial quantities of water, shipment costs associated with transport of micronutrient solutions are also avoided.

Moreover, the manufacturing methods and products of the present invention provide a micronutrient product which has a relatively high concentration of metal (approximately 20%-25% by weight), is economical to ship, requires less storage space than other metal chelates, and is easier to use in conjunction with other fertilizer materials. Storage problems associated with metal EDTA solutions, for example, crystallization, are minimized with the micronutrient products of the present invention. As described above, the partially chelated metal products produced in accordance with Examples I-IV described above do not appear to be susceptible to uncontrolled chelation, degradation or other reaction during storage at temperatures of up to 140°F for limited periods. Container disposal problems also associated with metal EDTA solutions are also avoided with the micronutrient products of the present invention. The preferred plastic bags used to package the products as described above in Examples I-IV, can be recycled or landfilled.

Presently preferred embodiments of the present invention and many of its improvements have been described with a degree of particularity. It should be understood that this description has been made by way of preferred examples, and that the invention is defined by the scope of the following claims.

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